

Dichalcogenide Bonding in Seven Alkali-Metal Actinide Chalcogenides of the KTh₂Se₆ Structure Type

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The solid-state compounds CsTh₂Se₆, Rb_{0.85}U_{1.74}S₆, RbU₂Se₆, TIU₂Se₆, Cs_{0.88}(La_{0.68}U_{1.32})Se₆, KNp₂Se₆, and CsNp₂Se₆ of the AA_nQ₆ family (A = alkali metal or Tl; An = Th, U, Np; Q = S, Se, Te) have been synthesized by high-temperature techniques. All seven crystallize in space group *I*m*m**m* of the orthorhombic system in the KTh₂Se₆ structure type. Evidence of long-range order and modulation were found in the X-ray diffraction patterns of TIU₂Se₆ and CsNp₂Se₆. A 4a × 4b supercell was found for TIU₂Se₆ whereas a 5a × 5b × 5c supercell was found for CsNp₂Se₆. All seven compounds exhibit Q–Q interactions and, depending on the radius ratio R_{An}/R_A , disorder of the A cation over two sites. The electrical conductivity of RbU₂Se₆, measured along [100], is 6×10^{-5} S cm⁻¹ at 298 K. The interatomic distances, including those in the modulated structure of TIU₂Se₆, and physical properties suggest the compounds may be formulated as containing tetravalent Th or U, but the formal oxidation state of Np in the modulated structure of CsNp₂Se₆ is less certain. The actinide contraction from Th to U to Np is apparent in the interatomic distances.

Introduction

Binary actinide chalcogenides, An_xQ_y (An = Th, U, Np; Q = S, Se, Te), display a variety of interesting electrical and magnetic properties.^{1–5} The insertion of an alkali metal into these compounds allows for the subtle manipulation of the structures, and consequently their physical properties. For this reason, extensive investigations of the ternary A/An/Q (A = alkali metal or Tl) system have been undertaken. The compounds known for the lighter alkali metals are A₂US₃ (A = Li, Na),⁶ which crystallize in space group *C*2/*m* of the monoclinic system. In these, as well in the isostructural A₂UTE₃ (A = K, Rb) compounds,⁷ there is no Q–Q bonding, so the compounds charge balance with U⁴⁺. The compound KUS₂,⁸ isostructural

to UOS,^{9,10} crystallizes in the space group *P*4/*nmm* of the tetragonal system. The lack of S–S bonding as well as magnetic measurements¹¹ suggest that unlike UOS it is a compound of U³⁺. The remainder of the A/An/Q compounds crystallize in orthorhombic space groups and feature some degree of Q–Q bonding. K₄USe₈¹² of the *F*dd2 space group is the only example of a molecular actinide species among the A/An/Q compounds. The U atom sits at the center of a distorted dodecahedron of four Se₂²⁻ ligands forming the [U(Se₂)₄]⁴⁻ complex anion that charge balances with four K⁺ cations. Once again, the U is assigned a formal oxidation state of +4. The compound CsUTE₆,¹³ belonging to the *Pnma* space group, has U atoms coordinated in tricapped trigonal-prismatic arrangements of Te atoms. If a Te–Te single-bond distance is arbitrarily taken to be less than 2.99 Å, then anionic chains in the [010] direction may be formulated as [U₂(Te₃)₃(Te₂)(Te)²⁻]. This also results in a formal oxidation state of +4 for U.

There are few well-characterized Np chalcogenides beyond the Np_xQ_y binaries. The known compounds include NpTQ

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- (1) Noël, H. *J. Less-Common Met.* **1986**, *121*, 265–270.
- (2) Shlyk, L.; Troc, R.; Kaczorowski, D. *J. Magn. Magn. Mater.* **1995**, *140*–*144*, 1435–1436.
- (3) Tougait, O.; Potel, M.; Noël, H. *J. Solid State Chem.* **1998**, *139*, 356–361.
- (4) Shlyk, L.; Troc, R. *Physica B* **1999**, *262*, 90–97.
- (5) Rossat-Mignod, J.; Burlet, P.; Fournier, J. M.; Pleska, E.; Quezel, S. *J. Nucl. Mater.* **1989**, *166*, 56–58.
- (6) Masuda, H.; Fujino, T.; Sato, N.; Yamada, K.; Wakushima, M. *J. Alloys Compd.* **1999**, *284*, 117–123.
- (7) Stöwe, K.; Appel-Colbus, S. Z. *Anorg. Allg. Chem.* **1999**, *625*, 1647–1651.
- (8) Padiou, J.; Lucas, J.; Prigent, J. C. *R. Acad. Sci. Ser. C* **1966**, *262*, 1869–1871.

- (9) Zachariasen, W. H. *Acta Crystallogr.* **1949**, *2*, 291–296.
- (10) Jin, G. B.; Raw, A. D.; Skanthakumar, S.; Haire, R. G.; Soderholm, L.; Ibers, J. A. *J. Solid State Chem.* **2010**, *183*, 547–550.
- (11) Padiou, J.; Lucas, J. *C. R. Acad. Sci. Ser. IIc: Chim.* **1966**, *263*, 71–73.
- (12) Sutorik, A. C.; Kanatzidis, M. G. *J. Am. Chem. Soc.* **1991**, *113*, 7754–7755.
- (13) Cody, J. A.; Ibers, J. A. *Inorg. Chem.* **1995**, *34*, 3165–3172.

($T = O, As, Sb; Q = S, Se, Te$),^{9,14–17} the superconducting Chevrel phase $Np_{1+x}Mo_6Se_8$,^{18,19} $NpCuSe_2$,²⁰ and the AMn_3S_3 ($A = K, Rb, Cs; M = Cu, Ag$) compounds.²¹ On the basis of magnetic measurements the compounds $Np-TQ$ ($T = As, Sb; Q = S, Se, Te$)^{14,16} and $Np_{1+x}Mo_6Se_8$ ^{18,19} contain Np^{3+} , but the compounds $NpOQ$ ($Q = S, Se$) contain Np^{4+} .^{22,23} Interatomic distances and structural comparisons in $NpCuSe_2$ indicate that the compound contains Np^{3+} ,²⁰ whereas those for the AMn_3S_3 compounds²¹ indicate Np^{4+} . Although, insofar as we know, there are no previously characterized ternary $A/Np/Q$ compounds, the Pu compound KPu_3Se_{8-x} is known.²⁴ This compound crystallizes in the $Cmcm$ space group isostructural to $RbDy_3Se_8$ ²⁵ and presumably contains Pu^{3+} .

The most studied members of the $A/An/Q$ ternary compounds belong to the AA_2Q_6 family. These can be further subdivided into two structure types, namely, $CsTh_2Te_6$ ²⁶ and KTh_2Se_6 ²⁷ (shown with their parent structure $ZrSe_3$ ²⁸ in Figure 1). The $CsTh_2Te_6$ structure type, which includes KTh_2Te_6 ²⁹ and $Tl_{1.12}U_2Te_6$,³⁰ crystallizes in space group $Cmcm$ of the orthorhombic system. The KTh_2Se_6 structure type, which includes $RbTh_2Se_6$,³¹ $K_{0.91}U_{1.79}S_6$,³² KU_2Se_6 ,²⁴ and CsU_2Se_6 ,²⁴ crystallizes in space group $Inmm$ of the orthorhombic system. Both structure types have AnQ_3 layers intercalated with A^+ cations. The difference between the two is that each successive AnQ_3 layer in the $Cmcm$ structure type is shifted by $1/2a$, whereas each successive AnQ_3 layer in the $Inmm$ structure type is shifted by $1/2(a + b)$.³²

Even though previous UV/vis, Raman,²⁷ diffuse reflectance,³² X-ray photoelectron spectroscopic,²⁴ electrical conductivity,³² and magnetic measurements^{27,30,32} suggest that the compounds in this family contain An^{4+} , the formula does not charge balance. Evidence of a $4a \times 4b$ supercell was found from electron diffraction measurements on the ATh_2Se_6 (K and Rb) compounds,²⁷ but the modulated structure was not solved. A

- (14) Charvillat, J. P.; Wojakowski, A.; Damien, D. *Proc. Int. Conf. Electron. Struct. Actinides*, 2nd **1976**, 469–473.
- (15) Blaise, A.; Charvillat, J. P.; Salmon, P.; Wojakowski, A. *Proc. Int. Conf. Electron. Struct. Actinides*, 2nd **1976**, 475–481.
- (16) Blaise, A.; Collard, J. M.; Fournier, J. M. *J. Phys., Lett.* **1984**, 45, L-571–L-576.
- (17) Wojakowski, A. *J. Less-Common Met.* **1985**, 107, 155–158.
- (18) Damien, D.; de Novion, C. H.; Gal, J. *Solid State Commun.* **1981**, 38, 437–440.
- (19) de Novion, C. H.; Damien, D.; Hubert, H. *J. Solid State Chem.* **1981**, 39, 360–367.
- (20) Wells, D. M.; Skanthakumar, S.; Soderholm, L.; Ibers, J. A. *Acta Crystallogr., Sect. E: Struct. Rep. Online* **2009**, 65, 114.
- (21) Wells, D. M.; Jin, G. B.; Skanthakumar, S.; Haire, R. G.; Soderholm, L.; Ibers, J. A. *Inorg. Chem.* **2009**, 48, 11513–11517.
- (22) Collard, J. M.; Blaise, A.; Bogé, M.; Bonnisseau, D.; Burlet, P.; Fournier, J. M.; Larroque, J.; Beauvy, M. *J. Less-Common Met.* **1986**, 121, 313–318.
- (23) Amoretti, G.; Blaise, A.; Bogé, M.; Bonnisseau, D.; Burlet, P.; Collard, J. M.; Fournier, J. M.; Quézel, S.; Rossat-Mignod, J.; Larroque, J. *J. Magn. Magn. Mater.* **1989**, 79, 207–224.
- (24) Chan, B. C.; Hulvey, Z.; Abney, K. D.; Dorhout, P. K. *Inorg. Chem.* **2004**, 43, 2453–2455.
- (25) Foran, B.; Lee, S.; Aronson, M. C. *Chem. Mater.* **1993**, 5, 974–978.
- (26) Cody, J. A.; Ibers, J. A. *Inorg. Chem.* **1996**, 35, 3836–3838.
- (27) Choi, K.-S.; Patschke, R.; Billinge, S. J. L.; Waner, M. J.; Dantus, M.; Kanatzidis, M. G. *J. Am. Chem. Soc.* **1998**, 120, 10706–10714.
- (28) Kröner, W.; Pfeith, K. *Z. Anorg. Allg. Chem.* **1965**, 336, 207–218.
- (29) Wu, E. J.; Pell, M. A.; Ibers, J. A. *J. Alloys Compd.* **1997**, 255, 106–109.
- (30) Tougaït, O.; Daoudi, A.; Potel, M.; Noël, H. *Mater. Res. Bull.* **1997**, 32, 1239–1245.
- (31) Schöllhorn, R. *Angew. Chem., Int. Ed. Engl.* **1996**, 35, 2338.
- (32) Mizoguchi, H.; Gray, D.; Huang, F. Q.; Ibers, J. A. *Inorg. Chem.* **2006**, 45, 3307–3311.

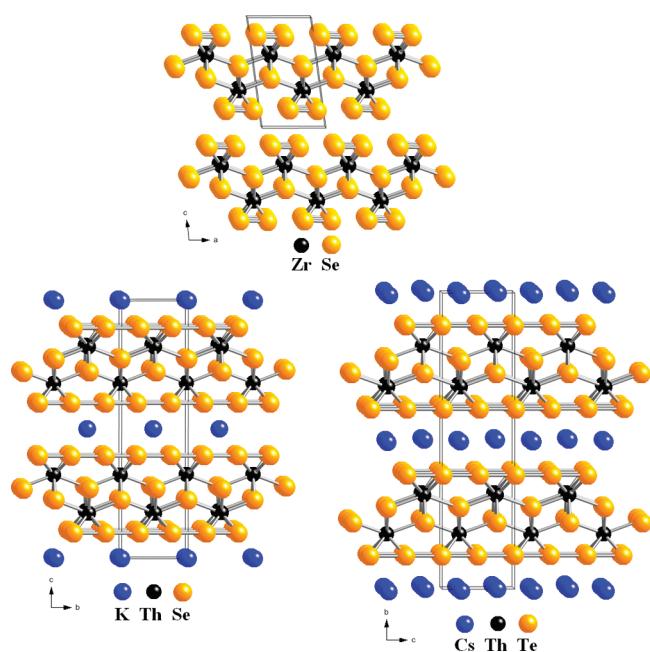


Figure 1. Comparison of the KTh_2Se_6 and $CsTh_2Te_6$ structure types along with their parent structure, $ZrSe_3$.

tentative model proposed is that one in every four Se_2^{2-} moieties is reduced to two Se^{2-} anions thus allowing for normal charge balance in the compound as $(A^+)_2(Th^{4+})_4(Se^{2-})_6(Se_2^{2-})_3$.²⁷

Here we report the syntheses of $CsTh_2Se_6$, $Rb_{0.85}U_{1.74}S_6$, Rb_2Se_6 , TIU_2Se_6 , $Cs_{0.88}(La_{0.68}U_{1.32})Se_6$, KNp_2Se_6 , and $CsNp_2Se_6$, their average structures, the modulated structures of TIU_2Se_6 and $CsNp_2Se_6$, and the electrical conductivity of RbU_2Se_6 .

Experimental Section

Syntheses. The following reagents were used as obtained from the manufacturers: K (Cerac, 98%), Rb (Strem, 99+%), Cs (Aldrich, 99.5%), La (Cerac, 99.9%), S (Mallinckrodt, 99.6%), Se (Cerac, 99.99%), Tl_2Se (Aldrich, 99.99%), $CsCl$ (Strem, 99.999%), Sb (Aldrich, 99.5%), $HgSe$ (Alfa Aesar, 99.9%), and Ag (Aldrich, 99.99+%).²³² Th foil was kindly supplied by Prof. M. G. Kanatzidis. Finely divided U powder was prepared from ^{238}U turnings (ORNL) by a modification³³ of the literature procedure.³⁴ Brittle ^{237}Np chunks (ORNL) were used as provided and crushed. The alkali-metal chalcogenide reactive fluxes³⁵ K_2Se , Rb_2S_3 , Rb_2Se_3 , and Cs_2Se_3 were prepared by the stoichiometric reactions of the alkali metals with S or Se in liquid NH_3 at 194 K.

Reactions were carried out in carbon-coated fused-silica tubes. The tubes were charged with reaction mixtures under an Ar atmosphere in a glovebox. The tubes were evacuated to 10^{-4} Torr and flame-sealed before being placed in a computer-controlled furnace. Selected single crystals from each reaction were examined by Energy Dispersive X-ray (EDX) analyses on a Hitachi S-3400 SEM. All the compounds are stable in air.

Synthesis of $CsTh_2Se_6$. $CsTh_2Se_6$ was synthesized by the reaction of Th (0.14 mmol), Cs_2Se_3 (0.14 mmol), and Se (0.39 mmol) in accordance with the literature procedure.²⁹

Synthesis of $Rb_{0.85}U_{1.74}S_6$. The compound $Rb_{0.85}U_{1.77}S_6$ was obtained from the reaction of U (0.5 mmol), Rb_2S_3 (0.95 mmol), and S (3.9 mmol). The sample was heated to 973 K in 15 h, kept at 973 K for 120 h, cooled at 3 K h^{-1} to 373 K, and then the

- (33) Bugaris, D. E.; Ibers, J. A. *J. Solid State Chem.* **2008**, 181, 3189–3193.
- (34) Haneveld, A. J. K.; Jellinek, F. *J. Less-Common Met.* **1969**, 18, 123–129.
- (35) Sunshine, S. A.; Kang, D.; Ibers, J. A. *J. Am. Chem. Soc.* **1987**, 109, 6202–6204.

furnace was turned off. The reaction mixture was washed free of flux with *N,N*-dimethylformamide and then dried with acetone. Black needles of $\text{Rb}_{0.85}\text{U}_{1.77}\text{S}_6$ were obtained in about 10 wt % yield based on U. EDX analysis of the crystals showed the presence of Rb, U, and S.

Synthesis of RbU_2Se_6 . The reaction mixture consisted of U (0.17 mmol), Rb_2Se_3 (0.17 mmol), Se (1.5 mmol), and Sb (0.17 mmol). The reaction mixture was placed in a furnace where it was heated to 923 K in 25 h, kept at 923 K for 120 h, cooled at 2 K h^{-1} to 423 K, and then quickly cooled to 298 K. The product was washed free of flux with *N,N*-dimethylformamide and then dried with acetone. It contained long golden black plates of RbU_2Se_6 in about 40 wt % yield based on U. EDX analysis of selected crystals showed the presence of Rb, U, and Se, but not of Sb.

Synthesis of TIU_2Se_6 . The reaction mixture consisted of U (0.17 mmol), HgSe (0.25 mmol), Ti_2Se (0.13 mmol), and Se (0.25 mmol). The mixture was placed in a furnace where it was heated to 1073 K in 24 h, kept at 1073 K for 7 days, cooled at 2 K h^{-1} to 479 K, and then the furnace was turned off. The product consisted of golden black plates of TIU_2Se_6 in about 20 wt % yield based on U, as well as black prisms of recrystallized HgSe. EDX analysis of selected crystals of TIU_2Se_6 showed the presence of Ti, U, and Se, but not of Hg.

Synthesis of $\text{Cs}_{0.88}(\text{La}_{0.68}\text{U}_{1.32})\text{Se}_6$. The reaction mixture consisted of U (0.13 mmol), La (0.13 mmol), Se (0.50 mmol), and CsCl (1.2 mmol). The mixture was placed in a furnace where it was heated to 1123 K in 17 h, kept at 1123 K for 6 days, cooled at 3 K h^{-1} to 673 K, and then the furnace was turned off. The product was washed free of salt with deionized water and then dried with acetone. It contained black plates of $\text{Cs}_{0.88}(\text{La}_{0.68}\text{U}_{1.32})\text{Se}_6$ in about 25 wt % yield based on U. EDX analysis of selected crystals showed the presence of Cs, La, U, and Se, but not of Cl.

Caution! ^{237}Np and any ingrown daughter products are α - and γ -emitting radioisotopes and as such are considered a health risk. Their use requires appropriate infrastructure and personnel trained in the handling of radioactive materials. Our procedures used for the syntheses of Np compounds have been described elsewhere.²¹

Synthesis of KNp_2Se_6 . The reaction mixture consisted of Np (0.08 mmol), K_2Se (0.03 mmol), and Se (0.24 mmol). It was placed in a furnace where it was heated to 1173 K in 15 h, kept at 1173 K for 96 h, cooled at 5.75 K h^{-1} to 603 K, and finally cooled at 3 K h^{-1} to 298 K. The product consisted of small black rectangular plates of KNp_2Se_6 in low yield. These were manually extracted from the product mixture.

Synthesis of CsNp_2Se_6 . The reaction mixture consisted of Np (0.09 mmol), Cs_2Se_3 (0.10 mmol), Se (0.46 mmol), and Ag (0.04 mmol). It was placed in a furnace where it was heated to 373 K in 4 h, kept at 373 K for 20 h, heated to 873 K in 6 h, kept at 873 K for 96 h, cooled at 4 K h^{-1} to 573, and then quickly cooled to room temperature. The product contained thick black rectangular plates of CsNp_2Se_6 in low yield; these were manually extracted from the product mixture.

Structure Determinations. Single-crystal X-ray diffraction data were collected with the use of graphite-monochromatized MoK α radiation ($\lambda = 0.71073 \text{ \AA}$) on a Bruker SMART-1000 CCD diffractometer³⁶ for the two Rb compounds and on a Bruker APEX2 CCD diffractometer³⁷ for the remaining five compounds. The crystal-to-detector distance was 5.023 cm for the Th and U compounds, and 5.106 cm for the Np compounds. Crystal decay was monitored by recollecting 50 additional frames at the end of the data collection. Data were collected by a scan of 0.3° in ω in groups of 606 frames at ϕ settings of $0^\circ, 90^\circ, 180^\circ$, and 270° . The collection

of intensity data was carried out with the program SMART³⁶ for the Rb compounds and with the APEX2³⁷ software for all other compounds. Cell refinement and data reduction were carried out with the use of the program SAINT v7.34a in APEX2.³⁷ Face-indexed absorption corrections were performed numerically with the use of the program SADABS.³⁸ Then that program was employed to make incident beam and decay corrections. The structures were solved with the direct-methods program SHELXS and refined with the least-squares program SHELXL.³⁹ Each final refinement included a secondary extinction correction. The program STRUCTURE TIDY⁴⁰ was used to standardize the positional parameters. Reciprocal lattice plots of data were investigated with the program RLATT,³⁷ and all modulated data were refined with the program JANA2006.⁴¹

With the use of reflections consistent with space group $Imm\bar{m}$ the structures were solved in a routine fashion, taking into consideration nonstoichiometry, cation disorder, or both as required. The site occupancies for Rb and U in $\text{Rb}_{0.85}\text{U}_{1.74}\text{S}_6$ were refined in the same manner as described previously for $\text{K}_{0.91}\text{U}_{1.79}\text{S}_6$.³² Additionally, for RbU_2Se_6 and the Cs compounds, the A cation was disordered over two sites, A(1) and A(2), as found for the Cs cation in the isostructural compound CsU_2Se_6 .²⁴ Independent refinements of these occupancies confirmed full occupancy. Consequently, in the final refinements the site occupancies were constrained to α for A(2) and $1 - \alpha$ for A(1). The refined values of α were found to be 0.135(7), 0.201(6), and 0.268(8) for RbU_2Se_6 , CsTh_2Se_6 , and CsNp_2Se_6 , respectively. The Cs was also disordered over two sites in $\text{Cs}_{0.88}(\text{La}_{0.68}\text{U}_{1.32})\text{Se}_6$. In addition, the La/U ratio was not known, as both La and U occupy the same crystallographic site. As a result that ratio along with independent site occupancies of Cs(1) and Cs(2) were refined. The occupancies of Cs(1) and Cs(2) refined to 0.703(9) and 0.176(8), respectively. Although the refinement of site occupancies from X-ray diffraction data to determine composition is often unreliable and is not a recommended procedure,⁴² in this instance the result of $\text{Cs}_{0.88(2)}\text{La}_{0.68(4)}\text{U}_{1.32(4)}\text{Se}_6$ makes chemical sense, corresponding as it does to an oxidation state of +3.9(1) for U if we assume an oxidation state of +3 for La.

Additional experimental details for conventional structure solutions are given in Table 1 and in the Supporting Information. Selected metrical details are presented in Table 2.

During the refinements of TIU_2Se_6 and CsNp_2Se_6 it became evident that a number of relatively weak observed reflections did not conform to space group $Imm\bar{m}$. Further examination of reciprocal lattice plots of these data revealed both the strong Bragg reflections used in the conventional refinement and weaker satellite reflections. Those satellite reflections occur at approximately $0.25a$ and $0.25b$ for TIU_2Se_6 and $0.4a$ and $0.2b$ for CsNp_2Se_6 . The modulated reflection data were integrated using initial q-vectors of $\mathbf{q} = 0.25\mathbf{a}^* + 0.25\mathbf{b}^*$ for TIU_2Se_6 and $\mathbf{q} = 0.4\mathbf{a}^* + 0.2\mathbf{b}^*$ for CsNp_2Se_6 . No satellite reflections were observed in the diffraction patterns of the other five compounds. It may be that evidence of such satellite reflections could be found at lower temperatures.

The data for TIU_2Se_6 were refined to a monoclinic unit cell $a = 4.0391(5)$, $b = 5.551(7)$, $c = 21.655(3)$, $\beta = 90.022(2)$, $\mathbf{q} = 0.2433(2)$ $\mathbf{a}^* + 1/4 \mathbf{b}^*$ in the super space group $P2/m(\alpha\beta\theta)00$ to an overall R index of 0.096 ($R_w = 0.229$) with 1221 main reflections (667 observed), 2402 first-order satellite reflections (258 observed), and 1720 second-order reflections (16 observed). During final refinement the modulation was constrained to a commensurate $4a \times 4b \times c$ supercell ($\mathbf{q} = 1/4\mathbf{a}^* + 1/4\mathbf{b}^*$) resulting in an overall R index of 0.096 ($R_w = 0.218$) for 198 parameters and 5343 observed reflections.

(36) SMART Version 5.054 Data Collection and SAINT-Plus Version 6.45a Data Processing Software for the SMART System; Bruker Analytical X-Ray Instruments, Inc.: Madison, WI, 2003.

(37) APEX2 Version 2009.5-1 and SAINT Version 7.34a Data Collection and Processing Software; Bruker Analytical X-Ray Instruments, Inc.: Madison, WI, 2009.

(38) Sheldrick, G. M. *SADABS*; University of Göttingen: Göttingen, Germany, 2008.

(39) Sheldrick, G. M. *Acta Crystallogr., Sect. A: Found. Crystallogr.* **2008**, *64*, 112–122.

(40) Gelato, L. M.; Parthé, E. *J. Appl. Crystallogr.* **1987**, *20*, 139–143.

(41) Petricek, V.; Dusek, M.; Palatinus, L. *Jana2006, The crystallographic computing system*; Institute of Physics: Praha, Czech Republic, 2006.

(42) Mironov, Y. V.; Cody, J. A.; Albrecht-Schmitt, T. E.; Ibers, J. A. *J. Am. Chem. Soc.* **1997**, *119*, 493–498.

Table 1. Crystal Data and Conventional Structure Refinements for AAn₂Q₆ Compounds^a

	CsTh ₂ Se ₆	Rb _{0.85} U _{1.74} S ₆	RbU ₂ Se ₆	TIU ₂ Se ₆	Cs _{0.88} (La _{0.68} U _{1.32})Se ₆	KNp ₂ Se ₆	CsNp ₂ Se ₆
fw	1070.75	686.32	1035.29	1154.19	999.38	986.86	1080.67
<i>a</i> (Å)	4.1777(4)	4.034(4)	4.066(3)	4.0387(5)	4.1540(5)	4.0377(7)	4.011(1)
<i>b</i> (Å)	5.6221(5)	5.427(5)	5.552(5)	5.5524(7)	5.6575(7)	5.525(1)	5.519(1)
<i>c</i> (Å)	24.015(2)	21.21(2)	22.44(2)	21.646(3)	23.721(3)	21.558(4)	24.060(6)
<i>V</i> (Å ³)	564.05(9)	464.2(7)	506.6(7)	485.4(1)	557.5(1)	480.9(2)	532.5(2)
<i>T</i> (K)	100(2)	153(2)	153(2)	100(2)	100(2)	296(2)	100(2)
ρ_c (g cm ⁻³)	6.305	4.910	6.786	7.897	5.954	6.815	6.739
μ (mm ⁻¹)	48.804	36.536	58.172	72.252	44.07	44.544	43.205
$R(F)^b$	0.029	0.026	0.029	0.034	0.042	0.029	0.044
$R_w(F_o^2)^c$	0.065	0.060	0.081	0.109	0.103	0.063	0.104
q^e	0.0361	0.03	0.04	0.04	0.03	0.0197	0.0299

^a For all structures, *Z* = 2, space group is *Immm*, and λ = 0.71073 Å. ^b $R(F) = \sum ||F_o| - |F_c|| / \sum |F_o|$ for $F_o^2 > 2\sigma(F_o^2)$. ^c $R_w(F_o^2) = \{\sum w(F_o^2 - F_c^2)^2 / \sum wF_o^4\}^{1/2}$. For $F_o^2 < 0$, $w^{-1} = \sigma^2(F_o^2)$; for $F_o^2 \geq 0$, $w^{-1} = \sigma^2(F_o^2) + (qF_o^2)^2$.

Table 2. Selected Interatomic Distances (Å) for AAn₂Q₆ Compounds^a

	CsTh ₂ Se ₆	Rb _{0.85} U _{1.74} S ₆	RbU ₂ Se ₆	TIU ₂ Se ₆	Cs _{0.88} (La _{0.68} U _{1.32})Se ₆	KNp ₂ Se ₆	CsNp ₂ Se ₆
A(1)-Q(2) × 8	3.767(1)	3.446(3)	3.490(2)	3.343(1)	3.780(1)	3.3478(7)	3.781(1)
A(2)-Q(2) × 8	3.725(1)		3.459(2)		3.624(1)		3.739(1)
An-Q(1) × 2	2.959(1)	2.775(3)	2.891(2)	2.884(1)	2.908(1)	2.8673(9)	2.869(2)
An-Q(1) × 2	3.0061(7)	2.823(2)	2.961(2)	2.9570(7)	2.9645(7)	2.9443(6)	2.9367(9)
An-Q(2) × 4	2.985(1)	2.847(2)	2.927(2)	2.922(1)	2.978(1)	2.9048(7)	2.898(1)
Q(2) ··· Q(2)	2.698(3)	2.106(5)	2.698(3)	2.704(3)	2.422(3)	2.681(2)	2.647(3)
	2.924(3)	3.321(4)	2.854(3)	2.849(2)	3.236(2)	2.844(2)	2.872(3)

^a The An and Q(1) atoms have crystallographic symmetry *mm2*, the Q(2) atom has crystallographic symmetry *m..*, and the A(1) and A(2) atoms have crystallographic symmetry *mmm*.

All refinements became unstable upon removal of the isotropic displacement parameter constraints. The weakness of the satellite reflections seriously degraded the refinements. Attempts to collect a more intense data set with the use of Cu radiation proved impossible because no suitable crystals could be found.

The modulated data for CsNp₂Se₆ were refined to an orthorhombic unit cell *a* = 4.0098(9) Å, *b* = 5.519(1) Å, *c* = 24.064(5) Å, \mathbf{q} = 0.396(3) \mathbf{a}^* + 0.1980(5) \mathbf{b}^* - 0.034(2) \mathbf{c}^* . When the orthorhombic cell was transformed to the reduced triclinic cell the refinement improved and the triclinic unit cell *a* = 4.0146(9) Å, *b* = 5.525(1), *c* = 12.520(3), α = 102.785(3), β = 99.228(3), γ = 89.977(3), \mathbf{q} = 0.3958(2) \mathbf{a}^* + 0.1966(3) \mathbf{b}^* + 0.2088(6) \mathbf{c}^* resulted. The incommensurate structure was solved in the triclinic superspace group *P1(αβγ)0* to an *R* index of 0.053 from 3565 main reflections (2865 observed) and 7240 first-order satellite reflections (982 observed) with all displacement parameters of like atoms constrained to be equivalent. In the final refinement the data were constrained to a commensurate 5*a* × 5*b* × 5*c* supercell with \mathbf{q} = 0.4 \mathbf{a}^* + 0.2 \mathbf{b}^* + 0.2 \mathbf{c}^* , constraints on the displacement parameters were removed, the instability factor was increased to 0.055, and the Cs(2) occupancy factor was varied. This led to an overall *R* index of 0.0491 for 274 parameters and 3847 observed reflections. Even though refinement of the modulated structure improved with constraint of the *q*-vector indices to rational fractions, the temperature dependence of the *q*-vector was not investigated. This would be necessary to establish the true nature of the modulation.

Additional experimental details are given in Table 3 and in the Supporting Information. Selected metrical details are presented in Table 4.

Electrical Conductivity Measurement. The electrical resistivity of a single crystal of RbU₂Se₆ was measured along [100] between 235 and 325 K utilizing a Fluke 73 Series II multimeter while the sample was submerged inside a Quantum Design PPMS instrument where the sample temperature was controlled. The crystal, approximately 2 mm in length, was mounted with two leads constructed of 15 μm diameter Cu wire and 8 μm diameter graphite fibers, and attached with Dow 4929N silver paint. Below 235 K the total resistance of the sample was beyond the limit (32 MΩ) of the multimeter.

Table 3. Crystal Data and Modulated Structure Refinements for TIU₂Se₆ and CsNp₂Se₆^a

	TIU ₂ Se ₆	CsNp ₂ Se ₆
superspace group	<i>P2/m(αβγ)000</i>	<i>P1(αβγ)0</i>
fw	1154.2	1080.8
<i>a</i> (Å)	4.0391(5)	4.0146(9)
<i>b</i> (Å)	5.5510(7)	5.525(1)
<i>c</i> (Å)	21.654(3)	12.520(3)
α (deg)	90	102.785(3)
β (deg)	90	99.228(3)
γ (deg)	90.022(2)	89.977(3)
<i>q</i> -vector	1/4 \mathbf{a}^* + 1/4 \mathbf{b}^*	2/5 \mathbf{a}^* + 1/5 \mathbf{b}^* + 1/5 \mathbf{c}^*
<i>V</i> (Å ³)	485.5(1)	267.1(1)
<i>Z</i>	2	1
ρ_c (g cm ⁻³)	7.8925	6.717
μ (mm ⁻¹)	72.202	43.07
overall $R(F)^b/R_w(F_o^2)^c$	0.096/0.218	0.049/0.130
main reflections $R(F)/R_w(F_o^2)$	0.091/0.214	0.044/0.118
1st order reflections $R(F)/R_w(F_o^2)$	0.119/0.303	0.097/0.197
$S(obs)/S(all)$	6.25/2.49	0.96/0.83
number of refined parameters	198	274
constraints	18	5

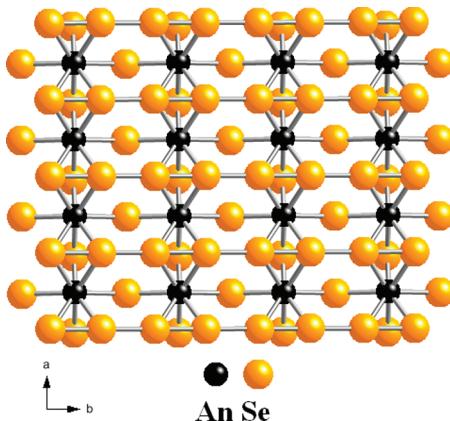
^a For both structures, *T* = 100(2) K and λ = 0.71073 Å. ^b $R(F) = \sum ||F_o| - |F_c|| / \sum |F_o|$ for $F_o^2 > 3\sigma(F_o^2)$. ^c $R_w(F_o^2) = \{\sum w(F_o^2 - F_c^2)^2 / \sum wF_o^4\}^{1/2}$. $w^{-1} = \sigma^2(F_o^2) + qF_o^2$; $q = 0.012$ for Cs and 0.0004 for TiL.

Results and Discussion

Syntheses. CsTh₂Se₆, Rb_{0.85}U_{1.74}S₆, and KNp₂Se₆ were prepared by the reactive flux method.³⁵ RbU₂Se₆ was synthesized inadvertently while attempting to obtain a quaternary compound in the Rb/U/Sb/Se system. Its synthesis proceeded from the reaction of U, Sb, Rb₂Se₃, and Se at 923 K with a 40 wt % yield. Attempts to produce RbU₂Se₆ rationally without Sb present were unsuccessful. It is possible that Sb reacted to form RbSb_xSe_y, *in situ*, to promote crystal growth, as was observed in the synthesis of RbTh₂Se₆.²⁷ TIU₂Se₆ was prepared in 20 wt % yield as the unintentional

Table 4. Selected Interatomic Distances (\AA) from the Modulated Structures of TIU_2Se_6 and CsNp_2Se_6

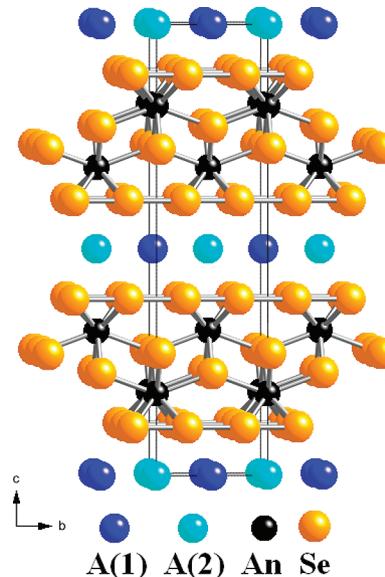
	TIU_2Se_6			CsNp_2Se_6		
	average	minimum	maximum	average	minimum	maximum
A(1)–Q(2)	3.34	3.21	3.45	3.79	3.57	4.02
A(2)–Q(2)	3.34	3.22	3.47	3.69	3.52	3.80
An–Q(1)	2.93	2.84	3.08	2.91	2.85	2.96
An–Q(2)	2.93	2.89	2.97	2.90	2.86	2.95
Q(2)–Q(2)	2.78	2.50	3.01	2.76	2.46	3.10

**Figure 2.** One-half of the $[An_2Se_6]^{2-}$ layer in the AAn_2Se_6 structure, as viewed down [001].

product from the reaction of U, HgSe, Ti_2Se , and Se at 1073 K. Attempts to produce TIU_2Se_6 rationally without HgSe were unsuccessful. $\text{Cs}_{0.88}\text{La}_{0.68}\text{U}_{1.32}\text{Se}_6$ was prepared from the reaction of U, La, Se, and CsCl at 1123 K. Interestingly, it is the first compound of the $K\text{Th}_2\text{Se}_6$ structure type²⁷ to be synthesized with an excess of an alkali-metal halide salt rather than with an alkali-metal polychalcogenide reactive flux. CsNp_2Se_6 was synthesized inadvertently while trying to obtain the Se analogue of CsAgNpS_3 .²¹ Owing to the scarcity of Np metal, a rational synthesis of the compound was not attempted.

Structures. The average structures of all seven compounds belong to the $K\text{Th}_2\text{Se}_6$ structure type²⁷ in space group $Immm$ of the orthorhombic system. These structures can be visualized by insertion of A atoms into the corresponding $An\text{Q}_3$ structure (Figure 1). ($\text{Cs}_{0.88}\text{La}_{0.68}\text{U}_{1.32}\text{Se}_6$ is analogous to the previously reported CsU_2Se_6 ,²⁴ with both La and U disordered over the same site.) In all the compounds, each An atom is coordinated by eight Q atoms in a distorted bicapped trigonal-prismatic geometry. The $An\text{Q}_8$ trigonal prisms share faces in the [100] direction and edges in the [010] direction, resulting in $[An_2Q_6]^{2-}$ layers (Figure 2). The A⁺ cations are intercalated between layers in either one or both of the two square-prismatic sites (Figure 3). Part of the coordination at these sites involves either four short or four long Q–Q interactions.

The An ionic radii decrease from Th to Np. The reduction in an An coordination sphere decreases the An–Se interatomic distances and consequently the Se–Se interactions. These reductions accordingly cause the $[An_2Se_6]^{2-}$ layer to contract. This contraction determines whether the A cation can occupy only one square-prismatic site or disorder over both of the square-prismatic sites. The occurrence of disorder depends on the radius ratio between the eight-coordinate actinide (R_{An}) and the alkali-metal (R_A) ions. A similar trend of the cation radius ratio determines whether members of the

**Figure 3.** Crystal structure of an AAn_2Se_6 compound showing the disorder over the A(1) and A(2) sites. View is down [100].**Table 5.** Radius Ratio versus A Site Occupancy in the AAn_2Se_6 Compounds

compound	R_{An}	R_A	R_{An}/R_A	ordered	A(2) occupancy		reference
					reference		
$\text{Rb}_{0.86}\text{Zr}_2\text{Se}_6$	0.84	1.61	0.52	N	0.36	44	
$\text{K}_{1.10}\text{Zr}_2\text{Se}_6$	0.84	1.51	0.56	N	0.46	44	
CsNp_2Se_6	0.98	1.74	0.56	N	0.268(8)	this work	
CsU_2Se_6	1.00	1.74	0.57	N	0.30	24	
CsTh_2Se_6	1.05	1.74	0.60	N	0.201(6)	this work	
$\text{Cs}_{0.88}(\text{La}_{0.68}-\text{U}_{1.32})\text{Se}_6$	1.05	1.74	0.60	N	0.176(8)	this work	
RbU_2Se_6	1.00	1.61	0.62	N	0.135(7)	this work	
TIU_2Se_6	1.00	1.59	0.63	Y		this work	
KNp_2Se_6	0.98	1.51	0.65	Y		this work	
RbTh_2Se_6	1.05	1.61	0.65	Y		27	
KU_2Se_6	1.00	1.51	0.66	Y		24	
KTh_2Se_6	1.05	1.51	0.70	Y		27	

ALnS_2 ($\text{Ln} = \text{La}–\text{Lu}$) family of compounds adopt the NaCl or the $\alpha\text{-NaFeO}_2$ structure type.⁴³ As shown in Table 5, which also includes the only nonactinide members of the $K\text{Th}_2\text{Se}_6$ structure type ($\text{K}_{1.10}\text{Zr}_2\text{Se}_6$ and $\text{Rb}_{0.86}\text{Zr}_2\text{Se}_6$),⁴⁴ compounds with smaller values of R_{An}/R_A exhibit disorder. The present compounds, RbU_2Se_6 and TIU_2Se_6 , are at the borderline between ordered and disordered structures. Note that the disorder is only found in the Se compounds and neither $\text{K}_{0.91}\text{U}_{1.79}\text{S}_6$ ³² nor $\text{Rb}_{0.85}\text{U}_{1.75}\text{S}_6$ exhibits disorder of the A cation.

(43) Bronger, W. In *Crystallography and Crystal Chemistry of Materials with Layered Structures*; Lévy, F., Ed.; D. Reidel Publishing Company: Dordrecht-Holland, 1976; pp 93–125.

(44) Klepp, K. O.; Harringer, N. A.; Kolb, A. Z. *Naturforsch. B: Chem. Sci.* **2002**, 57, 1265–1269.

All general cation–anion interatomic distances are normal. Thus: U–S, 2.775(3)–2.847(2) Å vs 2.645(2)–3.099(3) Å in Cu₂U₃S₇,⁴⁵ Th–Se, 2.959(1)–3.0061(7) Å vs 2.9309(7)–3.1611(7) Å in Mn₂ThSe₃,⁴⁶ U–Se, 2.884(1)–2.961(2) Å vs 2.850(2)–3.126(3) Å in MnUSe₃,⁴⁶ Rb–S, 3.446(3) Å vs 3.312(1)–3.483(1) Å in RbInS₂,⁴⁷ K–Se, 3.3478(7) Å vs 3.190(7)–3.694(8) Å in K₄USe₈,¹² Rb–Se, 3.459(2) and 3.490(2) Å vs 3.402(1)–3.600(2) Å in RbInSe₂,⁴⁷ Tl–Se, 3.343(1) Å vs 3.344(4)–3.403(4) Å in TiV₅Se₈,⁴⁸ and Cs–Se, 3.624(1)–3.781(1) Å vs 3.586(2)–3.825(1) Å in CsCuUSe₃.⁴⁹ The U(La)–Se distances of 2.908(1)–2.978(1) Å in Cs_{0.88}(La_{0.68}U_{1.32})Se₆ may also be compared to those of 3.0295(6) Å in RbLaSe₂.⁵⁰ Because of the disorder on the (La,U) site in Cs_{0.88}La_{0.68}U_{1.32}Se₆, the (La,U)–Se average interatomic distances are longer than a typical U–Se interatomic distance, but shorter than a typical La–Se interatomic distance. The Np–Se interatomic distances for KNp₂Se₆ and CsNp₂Se₆, which are in the range of 2.8673(9) to 2.9443(6) Å, are shorter than those found for Np³⁺ in NpCuSe₂ (2.9330(6) to 3.1419(6) Å).²⁰ Even though Nd (the 4f analogue of Np) does not form an isostructural compound, the Np–Se bonds here are much shorter than those found for 8-coordinate Nd in CsNd₅Se₈ (2.994(1) to 3.1616(5) Å)⁵¹ or 6-coordinate Nd in RbNdSe₂ (2.9591(6) to 3.4320(8) Å).⁵⁰ The Np³⁺–Se distance in NpCuSe₂²⁰ differs from that in NdCuSe₂⁵² by less than 0.021 Å. Clearly, the Np–Se distances in the present compounds are consistent with Np in the formal oxidation state +4, rather than +3.

The Q–Q interactions are of the utmost importance in understanding the KTh₂Se₆ structure type. Q(1) atoms are bonded only to the An atoms in the layers, but Q(2) atoms form infinite chains with alternating short and long Q–Q interactions. In Rb_{0.85}U_{1.75}Se₆, the short and long S–S interactions are 2.106(5) and 3.321(4) Å, respectively. The short interaction is indicative of an S–S single bond, whereas the long interaction is a non-bonding distance. Similarly, for Cs_{0.88}La_{0.68}U_{1.32}Se₆, the Se–Se interactions are 2.422(3) and 3.236(2) Å, a Se–Se single bond and a non-bonding distance, respectively. This may be contrasted with the similarly determined Se–Se interactions in the other five compounds where both the short and the long Se–Se interactions are significantly longer than Se–Se single bonds such as those observed in K₄USe₈ (2.385(4) and 2.401(4) Å).¹²

Modulated Structures. The data collected for TlU₂Se₆ display a two-dimensional 4a × 4b supercell like that found previously in KTh₂Se₆ and RbTh₂Se₆.²⁷ The Se–Se interatomic distances found here suggest the presence of Se₂²⁻ and Se²⁻ moieties and a range of Se–Se bond lengths from 2.50 to 3.01 Å. Even though a Se–Se bond length of 2.50 Å is long for a single bond, it is only

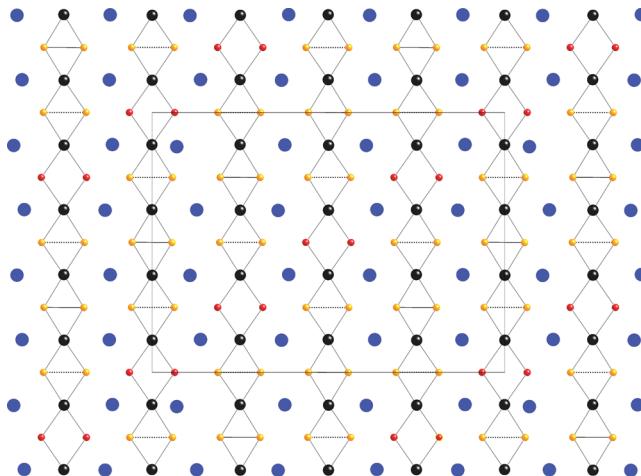


Figure 4. 4a × 4b Tl/U/Se layer from the supercell of TlU₂Se₆, as viewed down [001]. Red Se atoms are discrete Se²⁻ anions; orange Se atoms participate in Se–Se bonding. For Se–Se bonds: solid ≤ 2.5 Å, 2.5 < dashed ≤ 2.6 Å, and 2.6 < dotted ≤ 2.7 Å.

0.08 Å longer than those found in Cs_{0.88}La_{0.68}U_{1.32}Se₆. As with any structure solution, bond distance comparisons like those detailed above can allow for the formulation of an oxidation state model. Figure 4 displays a view down the [001] axis of the diselenide chains with neighboring Tl and U layers. A number of Se–Se interactions arbitrarily distinguished by their length are shown. The pattern of Se interactions seems to fit well with the suggested breaking of one out of four Se–Se bonds. The interatomic distances would likely change with a better refinement but the pattern of electron density should be maintained. As previously described,²⁷ this one-out-of-four model allows for an oxidation state description in this compound as (Tl⁺)₂(U⁴⁺)₄(Se²⁻)₆(Se₂²⁻)₃ if Se–Se interatomic distances longer than 2.7 Å are considered Se²⁻ anions.

In contrast to the 4a × 4b supercell found in the Th and U compounds, the present solution to the modulated structure of CsNp₂Se₆, solved in the reduced crystallographic triclinic unit cell, reflects a 5a × 5b × 5c supercell, but solutions in the orthorhombic cell suggest the supercell is also two-dimensional but 5a × 5b. The reflection conditions of the data suggest the symmetry is higher, possibly I2/m, but an accurate solution of a higher symmetry structure has not been found. The modulated structure of CsNp₂Se₆ also displays diselenide bonds shorter than those of the conventional solution with the shortest at 2.46 Å (Table 4). Because the Cs⁺ cations are also disordered over two positions, the modulated solution contains both charge density and site occupancy waves and Figure 5 displays a view down the [001] axis of CsNp₂Se₆ similar to Figure 4 for TlU₂Se₆. As may be expected from the electronic structure arguments presented below, the figure clearly shows the coupling between diselenide bond breaking and Cs occupancy.

Instead of the one-out-of-four model of diselenide bond breaking, the structure can be understood as the breaking of two out of every five diselenide bonds from the parent NpSe₃ structure and the formation of a “new” diselenide from two neighboring Se atoms. The result is two isolated Se²⁻ anions with Se–Se interatomic distances between 2.86 and 3.10 Å and essentially a one-in-five model. No simple description of the bonding in CsNp₂Se₆ results from 5a × 5b × 5c supercell, as displayed in Figure 5 and Table 4. The larger supercell

(45) Daoudi, A.; Lamire, M.; Levet, J. C.; Noël, H. *J. Solid State Chem.* **1996**, *123*, 331–336.

(46) Ijjaali, I.; Mitchell, K.; Huang, F. Q.; Ibers, J. A. *J. Solid State Chem.* **2004**, *177*, 257–261.

(47) Huang, F. Q.; Deng, B.; Ellis, D. E.; Ibers, J. A. *J. Solid State Chem.* **2005**, *178*, 2128–2132.

(48) Klepp, K.; Boller, H. *J. Solid State Chem.* **1983**, *48*, 388–395.

(49) Huang, F. Q.; Mitchell, K.; Ibers, J. A. *Inorg. Chem.* **2001**, *40*, 5123–5126.

(50) Deng, B.; Ellis, D. E.; Ibers, J. A. *Inorg. Chem.* **2002**, *41*, 5716–5720.

(51) Folchhardt, M.; Schleid, T. *Z. Kristallogr., Supplement Issue* **1997**, *12*, 125.

(52) Ijjaali, I.; Mitchell, K.; Ibers, J. A. *J. Solid State Chem.* **2004**, *177*, 760–764.

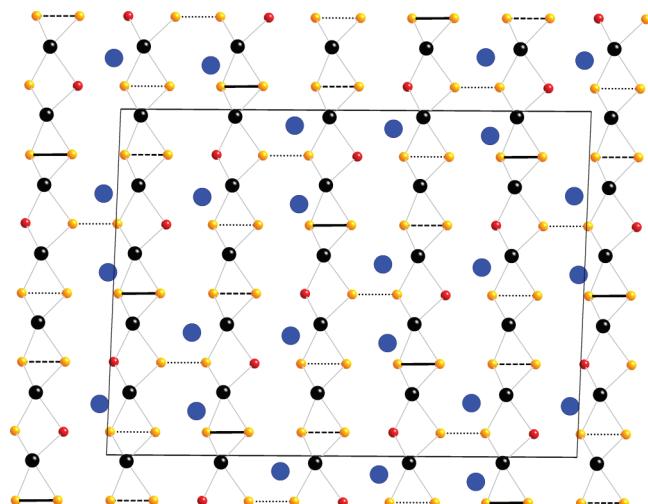


Figure 5. $5a \times 5b \times 5c$ Cs/Np/Se layer from the supercell of CsNp_2Se_6 , as viewed down [001]. Cs atoms are displayed on positions greater than 65% occupied. Red Se atoms are discrete Se^{2-} anions; orange Se atoms participate in Se—Se bonding. For Se—Se bonds: solid $\leq 2.5 \text{ \AA}$, $2.5 <$ dashed $\leq 2.6 \text{ \AA}$, and $2.6 <$ dotted $\leq 2.7 \text{ \AA}$.

found in CsNp_2Se_6 compared to that of Th and U could be an indication of the presence of Np^{3+} . If Se—Se bonds below 2.7 \AA are considered to be Se_2^{2-} then the formula can be written $\text{Cs}_5\text{Np}_{10}(\text{Se})_{14}(\text{Se}_2)_8$ and can be charge balanced with a one-to-ten $\text{Np}^{3+}/\text{Np}^{4+}$ ratio. If this were the case then localization of the Se_2^{2-} units might be expected, as was found with the trivalent La doping of CsU_2Se_6 to form $\text{Cs}_{0.88}\text{La}_{0.68}\text{U}_{1.32}\text{Se}_6$. The Np—Se interatomic distances in the modulated structure change very little from the conventional structure ranging from 2.85 to 2.96 \AA and still suggest Np^{4+} . A fully realized solution to the charge-balance problem may require quantification of the Cs vacancies by ICP measurements. If the Cs/Np ratio is reduced then the model also charge balances with Np^{4+} .

Electronic Structure. The electronic structure of the KTh_2Se_6 structure type has been explained by the insertion of an electron from the alkali metal into the S 3p σ^* or Se 4p σ^* orbital, respectively. For $\text{K}_{0.91}\text{U}_{1.79}\text{S}_6$ ³² it has been postulated that the formation of an infinite S—S chain with longer than normal bonding distances ($3.295(5) \text{ \AA}$) is less effective at decreasing the total energy of the system than is the formation of cationic vacancies. This rationalizes the non-stoichiometric A/U ratios for $\text{K}_{0.91}\text{U}_{1.79}\text{S}_6$ and $\text{Rb}_{0.85}\text{U}_{1.75}\text{S}_6$. From charge balance, these sulfur compounds can be formulated as $(\text{K}^+)^{0.91}(\text{U}^{3.96+})^{1.79}(\text{S}^{2-})_2(\text{S}_2^{2-})_2$ and $(\text{Rb}^+)^{0.85}(\text{U}^{4.09+})^{1.75}(\text{S}^{2-})_2(\text{S}_2^{2-})_2$ and hence as compounds of U^{4+} .

For the selenide compounds in the KTh_2Se_6 structure type, two alternative explanations of the electron distribution have been proposed. The first is that the electron from the alkali metal occupies one-fourth of the Se 4p σ^* orbitals.³² The second is that the donation of an electron from the alkali metal to the selenide chain breaks one of the diselenide bonds; this fits well with the modulated structures discussed above. The first explanation allows the formula of the selenides to be represented as $(\text{A}^+)(\text{An}^{4+})_2(\text{Se}^{1.25-})_4(\text{Se}^{2-})_2$ whereas the simplified model for the modulated Th and U structures can be written as $(\text{A}^+)_2(\text{An}^{4+})_4(\text{Se}^{2-})_6(\text{Se}_2^{2-})_3$. As described above, the $5a \times 5b \times 5c$ superstructure in CsNp_2Se_6 does not conform to this simplified model but the long-range order still displays shorter Se—Se bonds than are found in the

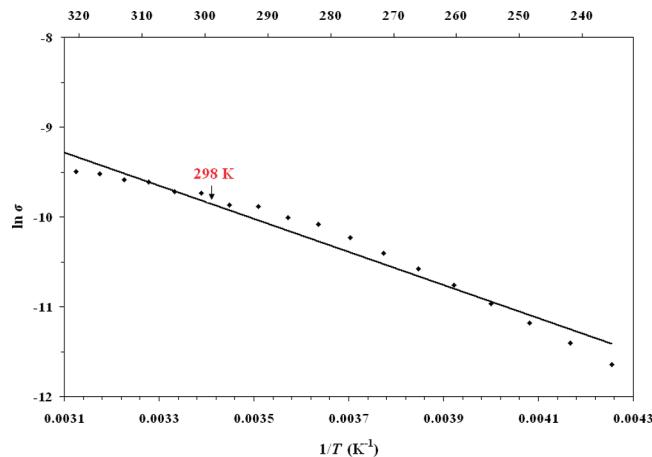


Figure 6. Plot of $\ln \sigma$ versus $1/T$ for RbU_2Se_6 , showing a linear Arrhenius fit to the data.

conventional structures. Currently the only Se exception is the disordered $\text{Cs}_{0.88}(\text{La}_{0.68}\text{U}_{1.32})\text{Se}_6$ compound where the extra electron from the alkali metal resides on the La^{3+} cation and the structure displays normal Se_2^{2-} bond distances. The formula for this doped compound can be represented as $(\text{Cs}^+)^{0.88}(\text{La}^{3+})^{0.68}(\text{U}^{3.85+})^{1.32}(\text{Se}^{2-})_2(\text{Se}_2^{2-})_2$. There seems to be a delicate balance between the oxidation/reduction potentials of the An and Se_2 moieties during the formation of these compounds. If the Np compounds could be synthesized from alkali halide salts rather than polyselenide fluxes, it might be possible to obtain Np compounds similar to the disordered lanthanide/actinide compound. It is unfortunate that the understanding of these potentials in solid-state environments remains rudimentary.

Electrical Conductivity of RbU_2Se_6 . Figure 6 displays the temperature dependence of the electrical conductivity of RbU_2Se_6 along [100]. At 298 K the conductivity σ is $6 \times 10^{-5} \text{ S cm}^{-1}$. Below 235 K, the sample resistance increased beyond the detection limits of the secondary multimeter ($32 \text{ M}\Omega$). This semiconductor shows a simple Arrhenius-type thermal activation $\sigma = \sigma_0 \exp(-E_a/(k_B T))$, where k_B is the Boltzmann constant. The calculated thermal activation energy, E_a , is $0.159(1) \text{ eV}$. To compare, the conductivity of KU_2Se_6 is $1 \times 10^{-3} \text{ S cm}^{-1}$, with a calculated activation energy of 0.27 eV .³² Other members of the AA_2Q_6 family, including ATH_2Se_6 and ATH_2Te_6 ($\text{A} = \text{K, Cs}$), have electrical conductivities below $1 \times 10^{-5} \text{ S cm}^{-1}$.^{26,29}

Final Remarks. The structural and physical property data collected here provide further support for an oxidation state of +4 for Th and U compounds in the KTh_2Se_6 structure type, but in the Np compounds the oxidation state is less clear. The Np—Se interatomic distances determined by the conventional refinement along with refinements of the modulated structures point toward Np^{4+} , but the $5a \times 5b \times 5c$ supercell for CsNp_2Se_6 cannot be charge balanced without mixed valent Np or reduced Cs occupancy. By setting an arbitrary limit on what can be described as an Se—Se single bond it is possible to present a simplified interpretation of the complex superstructures. It is interesting that the superspace refinements of the modulated structures of some LnTe_3 compounds⁵³ still lead to Te—Te distances far longer than

(53) Malliakas, C.; Billinge, S. J. L.; Kim, H. J.; Kanatzidis, M. G. *J. Am. Chem. Soc.* **2005**, *127*, 6510–6511.

the expected single-bond distance. With respect to the $\text{AA}\text{N}_2\text{Q}_6$ compounds, it is unfortunate that no previous attempts were made to solve and refine their modulated structures. In fact, to date few of the $\text{AA}\text{N}_2\text{Q}_6$ compounds have displayed measurable satellite data indicative of modulation. Further physical property measurements need to be made on the Np compounds to substantiate the +4 oxidation state as these would be the first examples insofar as we know of Np^{4+} in a selenide environment. Previous attempts to synthesize the Se analogues to Np^{4+} sulfides proved unsuccessful,²¹ and the other non-oxide selenides are considered Np^{3+} compounds.^{9,14–20} It is interesting that the Np compounds crystallize here in the $\text{AA}\text{N}_2\text{Q}_6$ structure that is stable for Th^{4+} and U^{4+} rather than in the RbDy_3Se_8 structure²⁵ known to be stable for Pu^{3+} .²⁴

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Supporting Information Available: The crystallographic file in CIF format for the conventional refinements of CsTh_2Se_6 , $\text{Rb}_{0.85}\text{U}_{1.74}\text{S}_6$, RbU_2Se_6 , TIU_2Se_6 , $\text{Cs}_{0.88}(\text{La}_{0.68}\text{U}_{1.32})\text{Se}_6$, KNp_2Se_6 , and CsNp_2Se_6 and a pdf file for the refinements of the modulated structures of TIU_2Se_6 and CsNp_2Se_6 . This material is available free of charge via the Internet at <http://pubs.acs.org>.